CORRECTION

In the article "Thermodynamic Functions of 1,2-*trans*-Difluorodiazine" by Emile Rutner [J. Chem. Eng. Data **7**, 398 (1962)], the following correction should be noted.

The column headings S^0/R and $-(F^0 - E_0^0)/RT$ should be transposed. The values for $-(F^0 - E_0^0)/RT$ which have been corrected are given below:

Table I. (Corrected)

T° K.	$-(F^0 - E_0^0)/RT^a$	T° K.	$-(F0 - E_0^0)/RT^a$
100	21.79	3100	42.50
200	24.61	3200	42.78
300	26.38	3300	43.05
400	27.77	3400	43.32
500	28.96	3500	43.58
600	30.01	3600	43.84
700	30.95	3700	44.08
800	31.81	3800	44.33
900	32.61	3900	44.56
1000	33.34	4000	44.79
1100	34.03	4100	45.02
1200	34.67	4200	45.24
1300	35.27	4300	45.45
1400	35.84	4400	45.66
1500	36.38	4500	45.87
1600	36.89	4600	46.07
1700	37.38	4700	46.27
1800	37.85	4800	46.47
1900	38.29	4900	46.66
2000	38.72	5000	46.84
2100	39.13	5100	47.03
2200	39.52	5200	47.21
2300	39.90	5300	47.39
2400	40.26	5400	47.56
2500	40.61	5500	47.73
2600	40.95	5600	47.90
2700	41.28	5700	48.07
2800	41.60	5800	48.23
2900	41.91	5900	48.39
3000	42.21	6000	48.55

[a] Dimensionless.

CORRECTION

In the article "Halogenated Phthalimidoacetic Acids" by L. M. Caswell and E. D. Martinez [J. Chem. Eng. Data **13**, 286-7 (1968)], the last sentence of the text should read: Table II summarizes data for the bonds corresponding to the phthalimidoacetic acid bonds at 292.5 nanometers (log ϵ_{max} 3.38) in 95% ethanol and at 269 nanometers (log ϵ_{max} 3.04) in dilute sodium hydroxide (*3*).

CORRECTION

In the article "Heats of Combustion of *trans*-Stilbene and *trans*-2,2′,4,4′,6,6′-Hexanitrostilbene (HNS)" by Sidney Marantz and George T. Armstrong [J. Chem. Eng. Data **13**, 118 (1968)], by mistake the heats of formation derived for the two compounds were given with reversed signs. When properly calculated from the experimental heats of combustion and auxiliary data given, ΔH_f° [*trans*-Stilbene (*c*)] is +32.68 kcal. mole^{-1} and ΔH_f° [HNS(*c*)] is +13.88 kcal. mole^{-1}. The heats of formation of the corresponding gases become +53.36 and +56.98 kcal. mole^{-1}, respectively.

For the calculation of the effect on $E(C=C)$ of the substitution of six nitro groups on stilbene, Domalski suggests (*2*), on the basis of a review of the data on TNT which he has made as part of a comprehensive review of the combustion calorimetry of organic compounds, that a careful recalculation of the data of Badoche (*1*) leads to ΔH_f° [TNT(*g*)] = +12.92 kcal. mole^{-1}, rather than +15.5 kcal. mole^{-1} which the authors used. The combustion data of Badoche lead to ΔH_f° [TNT(*c*)] = −15.38 kcal. mole^{-1}, and the sublimation data of Edwards (*3*) lead to ΔH_s [TNT(*c*)] = +28.3 kcal. mole^{-1}, from which the above value is obtained.

When applied to the calculation of the bond energy changes as in the Marantz and Armstrong article, these values lead to $\Delta E(C=C)$ = −1.7 kcal. mole^{-1} for the effect of six nitro groups substituted on *trans*-stilbene. Thus the weakening effect is much less pronounced than had been previously concluded.

A substantial uncertainty is attached to this value owing to differences in auxiliary data available. For instance from their own vapor pressure data, Nitta *et al.* (*4*) calculated ΔH_s [TNT(*c*)] = +24.43 kcal. mole^{-1}, which differs from Edwards by 3.87 kcal. mole^{-1}. If this were taken as the vaporization energy of TNT(*c*), then $\Delta E(C=C)$ would be −9.44 kcal. mole^{-1}.

The authors wish to thank S. Benson for calling attention to the error mentioned above and to E. S. Domalski for the new information given.

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(1) Badoche, M., *Bull. Soc. Chim.* **6**, 570 (1939).

(2) Domalski, E. S., National Bureau of Standards, Washington, D. C., private communication, 1968.

(3) Edwards, G., *Trans. Faraday Soc.* **46**, 423 (1950).

(4) Nitta, I., Seki, S., Momotani, M., Sato, K., *Nippon Kagaku Zasshi* **71**, 378-82 (1950).